

24-3 Acids, Bases, and Salts (764-770)

2 Types of Solutions - Acids and Bases

Acid: substance that contains at least one hydrogen ion that can be removed when dissolved in water \downarrow H^+

How do we show hydrogen as an ion?

Common Properties of Acids:

1. Tastes sour
2. Corrosive
3. Turns red for indicator

Common Household Acids

1. Citric acids -- citrus fruits
2. Lactic acid -- yogurt, butter
3. Acetic acid -- vinegar
4. Acetylsalicylic acid -- aspirin
5. Carbonic acid -- carbonated drinks
6. Nitric acid -- fertilizers

Base: substance that forms hydroxide ions OH^- in a water solution; accepts H^+ from acids

Properties

1. Crystalline Structure
2. Feel slippery
3. Bitter taste
4. Corrosive like acids
5. Turns blue for indicator

Common Household Bases

1. Egg Whites
2. Baking Powder
3. Milk of Magnesia and Antacids
4. Soap
5. Cleaners with ammonia
6. Fertilizers

Indicator: organic compound that changes color in acids and bases

What happens when acids and bases interact?

H^+ OH^-
OH⁻ and H⁺ come together and form H₂O } NEUTRALIZATION
usually form salts too



Salt: compound formed when negatives ions of acid combine with positive ions from base
-salt, baking soda, chalk